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The Borane Complex. First off it is very important to understand little bit about the structure and the properties of the borane molecule. Borane exists naturally as a very toxic gas and it exists as dimer of the general formula  $B_2H_6$  (diborane). Additionally, the dimer  $B_2H_6$  ignites spontaneously in air. Borane is commercially available in ether and tetrahydrofuran (THF), in these ...

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10.8.1. Anti-Markovnikov addition of HBr to alkenes We saw in section 10.4 that under normal conditions, HBr adds to an unsymmetrical alkene to form an alkyl halide where the H goes onto the less substituted carbon, and the Br goes onto the more substituted carbon – thus, it obeys Markovnikov's Rule. However, when heated

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in the presence of a dialkyl peroxide (often written as ROOR), a ...

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Hydrogen iodide (H I) is a diatomic molecule and hydrogen halide. Aqueous solutions of HI are known as hydroiodic acid or hydriodic acid, a strong acid. Hydrogen iodide and hydroiodic acid are, however, different in that

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the former is a gas under standard conditions, whereas the other is an aqueous solution of the gas.

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Acetoacetyl CoA is the precursor of HMG-CoA in the mevalonate pathway, which is essential for cholesterol biosynthesis. It also takes a similar role in the ketone bodies synthesis (ketogenesis) pathway of the liver. In the ketone bodies digestion pathway (in the tissue), it is no longer associated with having HMG-CoA as a product or as a reactant.. It is created from acetyl-CoA, a thioester ...

[????????????????](#)

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Alkohole (arabisch ??????, DMG al-ku??l) sind organische chemische Verbindungen, die eine oder mehrere an unterschiedliche aliphatische Kohlenstoffatome gebundene Hydroxygruppen (–O–H) besitzen.. Der Unterschied zwischen Alkoholen und anderen Verbindungen mit OH-Gruppen (z. B. Enole, Halbacetale oder Carbonsäuren) als Teil der funktionellen Gruppe ist, dass in Alkoholen jedes ...

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Teoria rezonansu chemicznego. Struktura Kekulégo nie tłumaczyła jednak dlaczego benzen i inne związki aromatyczne nie posiadają właściwości charakterystycznych dla węglowodorów nienasyconych, mimo 6 elektronów  $\pi$  w pierścieniu (sekszet

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elektronowy). Ponadto w miarę gromadzenia  
się materiału eksperymentalnego okazywało  
się, że wszystkie wiązania C-C w pierścieniu  
benzenowym ...

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